

MAD ORG. CHEM. MIN. #5

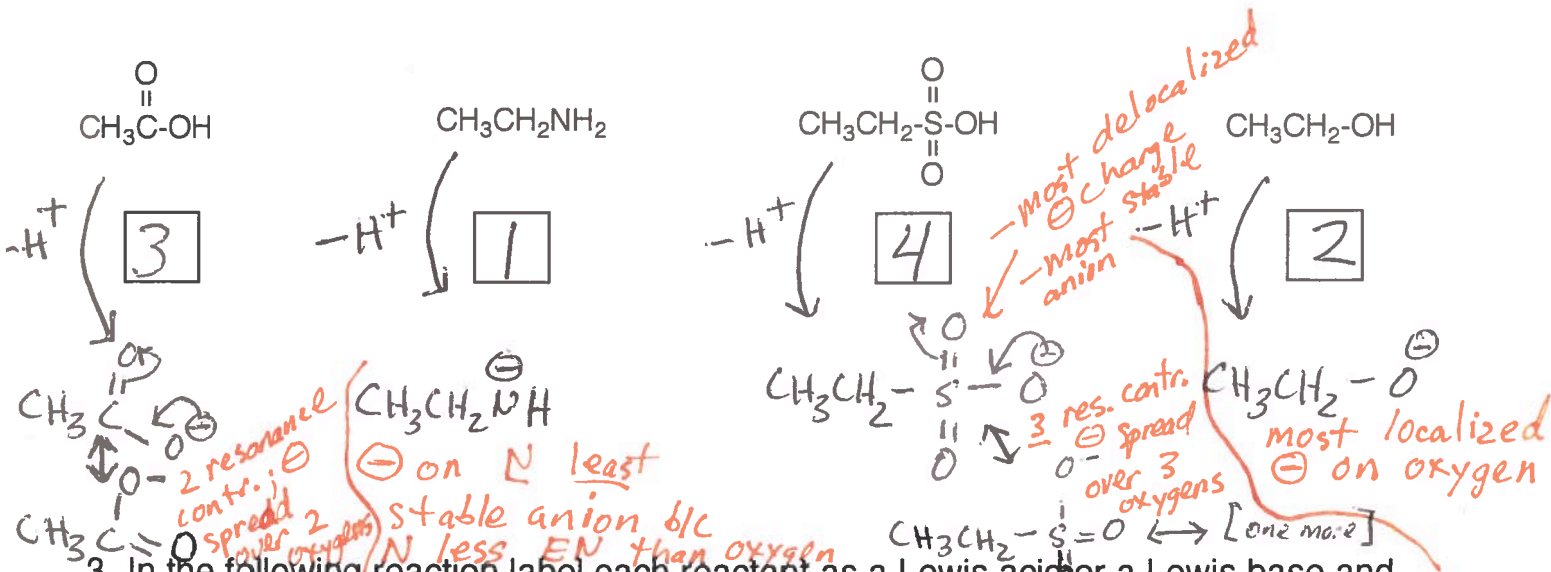
LAST NAME \_\_\_\_\_ FIRST NAME \_\_\_\_\_

IO # \_\_\_\_\_ CIRCLE CLASS TIME: 10AM 1:00 PM 4:00 PM  
*doesn't matter!*

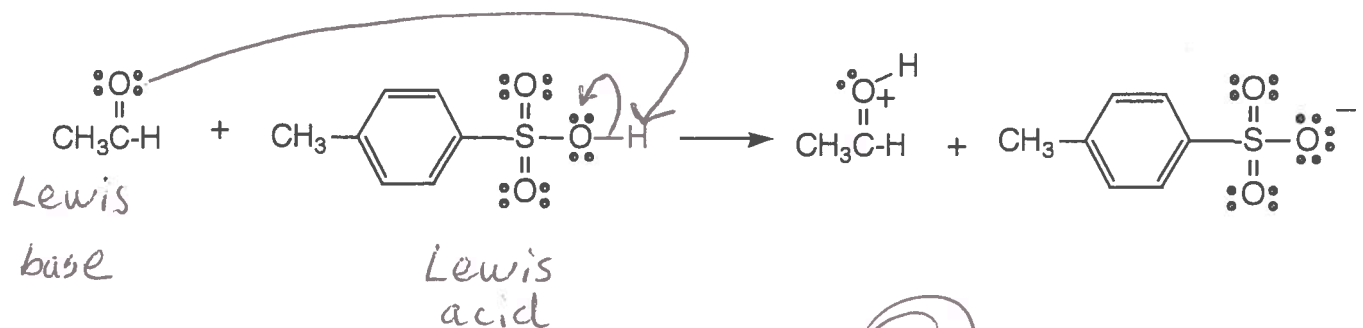
1. Which is the stronger acid, HF or HBr (EN of F = 4.0, EN of Br = 3.0)?

$F^-$  vs  $Br^-$  ← larger anion  
 more stable anion

2. Place the following compounds in order of increasing acidity (1=least acidic, 4=most acidic).



3. In the following reaction label each reactant as a Lewis acid or a Lewis base and provide curved arrows to indicate bond making and bond breaking (i.e., provide the mechanism). Be sure to indicate all movements of electrons with curved arrows.



4. Does the following equilibrium lie to the left or the right?

